

# THE FIRST CHEMICAL IDENTIFICATION OF POLYOLEFIN (PO) WASTE BLENDS USING INFRARED SPECTROSCOPY

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## ABSTRACT

Locally, the State of Kuwait is in need of good practices and national strategies that can take advantage of the accumulated solid waste, ridding the environment of various issues and stressors. In this work, compounded rolls of the various samples originating from plastic solid waste were used for the preparation of the test specimens. Samples were compounded to have the following percentages of virgin linear low density polyethylene and plastic film waste (wt./wt.%): 100/0, 75/25, 50/50, 25/75 and 0/100. The samples were compounded using blown film extrusion with a die head temperature of 175°C. Films of the different formulations were exposed to accelerated weathering in accordance with test method ASTM D 4329. Samples were mounted on the racks facing the ultra violet lamps with no empty spaces in the panels. This is in order to maintain a uniform repeatable test condition in accordance with cycle A procedure. The chemical printing of the evolved chemicals was achieved using Fourier Transform Infrared Spectroscopy and the spectra were recorded with a resolution of 4 cm<sup>-1</sup> using Diamond Universal ATR attachment. Four interferogram scans were averaged to give the spectra from 500 to 5000 cm<sup>-1</sup> using the Attenuated Total Reflection technique. A noted reduction in the carbonyl index determined from the FTIR spectra of the 75/25 samples supports the claim of the relative stability of the blend. This can lead to the development of a product from waste plastic material that reduced the burden of accumulation in developing and developed societies alike. In addition, the thermal properties determined along with the degradation kinetics parameters estimation revealed after the execution of this research project, have shown the stability of the blend for further development as a standalone product.

*Keywords:* plastic waste, polymer, solid waste, spectroscopy, carbonyl index, polyolefin.

## 1 INTRODUCTION

Thermoplastic polyolefin (PO) polymers such as polyethylene (PE) and polypropylene (PP), exhibit good mechanical and physical properties making them ideal for versatile use and applications [1]–[3]. These include packaging, consumer product commodities, leisure articles and building materials. The increasing demand on PO polymers has had its consequences as of late, namely in increasing petrochemical ventures around the globe and the increase in plastic solid waste (PSW) generation and accumulation. This is especially true when it comes to the State of Kuwait, where crude oil refining and petrochemical ventures are supported by the government of the state as it is considered the main contributor to the national gross domestic product (GDP) [4]. In addition, feedstock to plastic production industries is noted to be cheap due to the abundance in petroleum refinery cuts.

Various schemes make use of PSW, the form of mechanical recycling as a swift means for reducing the solid waste (SW) accumulation in communities. Mechanical recycling of plastics depends solely on the integrity of the produced article's integrity [5]. A mechanical recycling processing line is deemed successful when the used blend of virgin and PSW materials result in a product that substitute market accepted ones [6]. To this end, it is deemed essential to be able to investigate the properties of such blends using various chemical and thermal characterisation techniques.



Fourier transform infrared spectroscopy (FTIR) is a technique used in material characterization by examining the existing bonds in a sample. It works by searing different beams of varying frequencies and measuring the absorption of the beam by the sample. Each bond in a sample has a different energy and angle, thus absorbs the beam at specific wavelengths. The raw data collected from the different beams is modelled using Fourier Transform to present the measurement as absorption per wavelength. This allows the measurements of a wider range of wavelengths with high resolution. Moreover, weathering and ageing tests are typically conducted to estimate the durability of samples and study the change in the polymers matrix. Turku and Kärki [7] studied fire-retarded (aluminium trihydrate, zinc borate, melamine, graphite, titanium dioxide) wood-polypropylene composites (WPC) under accelerated conditions.  $\alpha$ -keto carbonyl groups of cellulose increased together with vinyl groups due to PP oxidation due to weathering. Moreover, the melamine peak was decreased on the front and backside of WPC due to leaching by water to the samples studied. Sharma et al. [8] characterised dumbbell specimens of Sago starch filled LLDPE composites and used infrared to study them. It was noted that ether and carbonyl peaks increased, suggesting the degradation of elastomer free radical reactions. Rajakumar et al. [9] used FTIR to determine the extent of degradation post natural weathering of PP samples. The carbonyl index increased slowly up to 40 days of summer exposure. Lactone, ester, ketone, and carboxylic acid indices showed a similar trend in summer and winter seasons.

Henceforth, this work was initiated to characterise various blends of PSW (originating from plastic film waste) with virgin commercial grade plastic (i.e. LLDPE) typically used in plastic film conversion. The work conducted in this study is a part of a continuing effort that stems out experimental results with the aim of having a comprehensive and novel picture of plastic film products rich in waste content [10]–[15]. The materials chosen are commercially available and used for various application, hence are of interest to study post use as PSW and converted plastics. This specific study focuses on the alterations caused by accelerated weathering (ageing) on the film samples originating from reclaimed PSW in Kuwait. The work carried out is useful for developing more resistant materials available on the market, as well as, extend the durability of the developed blends. Chemical alterations are reported here, in addition to the blends thermal stability post accelerated weathering. This paper will detail the efforts carried out and show the work conducted in detail with analysis of data gathered following on previous communications [16], [17]. To the best of author's knowledge, no such attempt was conducted in the past in investigating PSW blends.

## 2 MATERIALS AND METHODS

### 2.1 Samples compounding and preparation

Rolls of the various samples used in this work, were used for the preparation of the test specimens. Samples were compounded to have the following percentages of virgin linear low density polyethylene (LLDPE) and plastic film waste (virgin/waste, wt./wt.%): 100/0, 75/25, 50/50, 25/75 and 0/100. Throughout this manuscript the virgin LLDPE will be denoted as V and the PSW will be denoted as W. The samples were compounded using blown film extrusion with a die head temperature (DHT) of 175°C. Readers are referred to Al-Salem et al. [15]–[17] for further details on compounding and the PSW constituting elements as they are out of the scope in this communication. The thin rolls (100  $\mu$ m) were cut into standard test specimens size of 20 x 1 cm.



## 2.2 Accelerated weathering

Accelerated (artificial) indoor weathering (ageing) tests are commonly used for studying materials integrity. Films of the different formulations were exposed to AW in accordance with ASTM D 4329 [18]. The conducted weathering test was accomplished in the WMRU Laboratory in ELSRC. Samples were mounted on the racks facing the UV lamps with no empty spaces in the panels. This is in order to maintain a uniform repeatable test conditions. Cycle A procedure was used for general applications durability testing, i.e. 8 h of UV exposure at 60°C followed by 4 h of condensation at 50°C. At the end of each continuous weathering test, the chamber was cooled to room temperature and trays were set to rest on a flat surface for a minimum of 24 h. Samples were laid to rest for a minimum of 72 hours before characterization following the methodology of Johnson et al. [19], Al-Salem [20] and Al-Salem et al. [16]. A minimum of four replicates were exposed to the different exposure duration in the QUV machine chamber. Ultra Violet (UV) lamps irradiance was also selected according to ASTM D 4329 [18], and the lamp type was set to be 0.68 W m<sup>-2</sup> (irradiance) for normal lamp operation, which will maintain almost 5000 h of operation for the lamps. The irradiance sensor was calibrated every 400 h of lamp operation during the UV cycle under normal test temperature. The equipment used was cleaned every 800 hrs to remove scale deposits resulting from water evaporation during the condensation cycles.

## 2.3 FTIR analysis

Attenuated Total Reflection (ATR) works by producing the infrared (IR) beam into a solid crystal with a high refractive index. The sample is typically placed on the crystal and firmly pressed on it. The IR beam will reflect at the surface of the crystal creating evanescent waves in the sample. The IR beam reaches a detector at the other side of the crystal and the measurement is determined based on the absorption of the evanescent beam allowing the desired measurement to be conducted. The equipment used in this work was a Perkin Elmer UK (Spectrum Two Model) FTIR-spectrometer. Spectra were recorded with a resolution of 4 cm<sup>-1</sup> using Diamond Universal ATR attachment; four interferogram scans were averaged to give the spectra from 500 to 5000 cm<sup>-1</sup> using the Attenuated Total Reflection (ATR) technique. At least two replicas of samples were used to recorded each spectra. Figures below show the equipment used.

## 2.4 Thermal characterisation

Thermal degradation of virgin/waste blend samples was investigated using a Shimadzu TGA-50 thermobalance equipped with a data acquisition/analysis software (TA Instrument) set to record the data every second under five heating rates ( $\beta$ ) (i.e. 5, 10, 15, 20 and 25°C min<sup>-1</sup>). A constant flow of pure (99.99%) dry nitrogen with a flow rate of 50 ml min<sup>-1</sup> was maintained throughout the experiments. The measurements were conducted using 5±0.1 mg samples from room temperature to 550°C made with triplicates showing high repeatability with standard deviation not exceeding 1% in accordance with the International Confederation for Thermal Analysis and Calorimetry (ICTAC) recommendations previously published in Vyazovkin et al. [21], [22] for non-isothermal (dynamic) thermogravimetry; to diminish sample size influences on the kinetics [21]–[23]. Furthermore, samples of 1±0.01 mg were subjected to thermogravimetry in the same instrument under a constant flow of pure oxygen gas (99.99%) with a flow rate of 50 ml min<sup>-1</sup> ( $\beta$  equal to 15°C min<sup>-1</sup>). A Perkin Elmer (Model Jade, US) equipped with PYRIS analysis software was used to test control and exposed



samples using a  $5 \pm 0.1$  mg samples taken from the middle section of the specimens. Aluminium oxide ( $\text{Al}_2\text{O}_3$ ) crucibles made were used for both samples and reference materials. Crystallinity measurements were determined using scans of the second heating cycle between 50 to 230°C based on the peak area of the heat flow curve between 60°C and 130°C, with a nitrogen ( $\text{N}_2$ ) gas flowrate of 20 ml  $\text{min}^{-1}$  and a heating rate of 10°C  $\text{min}^{-1}$ . Cooling rate was set at 15°C  $\text{min}^{-1}$  in similar conditions [24], [25]. The second heating cycle was used in the crystallinity analysis following the initial heating run, as it will eliminate inherited effects of thermal histories and weathering on the specimens [25]. Oxidation induction temperature (OIT) was conducted using similar weight for the studied samples. Samples were heated using a rate of 10°C  $\text{min}^{-1}$  with a  $\text{N}_2$  flowrate of 20 ml  $\text{min}^{-1}$ . Samples were then held for 5 minutes at 70°C to settle followed by the introduction of oxygen ( $\text{O}_2$ ) gas flowrate of 20 ml  $\text{min}^{-1}$ . Samples were then heated from 70 to 230°C at a rate of 2°C  $\text{min}^{-1}$  [26].

### 3 RESULTS AND DISCUSSION

Accelerated weathered PO polymer blends originating from PSW blends, were analyzed in order to characterize and identify the chemical degradation products. The weathering extents were conducted in accordance with the experimental procedure described earlier and to the extent of the threshold limit of each formulation and described earlier in [16]. The carbonyl group, which is formed by the oxidation of polyethylene [26]–[28], was quantified using carbonyl index (CI) estimations based on the data obtained from the FTIR analysis (eqn (1)). Other chemical products noted to be prominent in polymer degradation were also analysed (e.g. lactone, ester/aldehydes, and carboxylic acid indices). This was done to monitor the extent of degradation by dividing the corresponding peak to the reference polyethylene (PE) band (eqns (2)–(4)).

$$\text{Carbonyl Index (CI)} = \frac{\text{Absorbance at } 1714 \text{ cm}^{-1}}{\text{Absorbance at } 1470 \text{ cm}^{-1}(\text{PE Band})}, \quad (1)$$

$$\text{Lactone Index} = \frac{\text{Absorbance at } 1770 \text{ cm}^{-1}}{\text{Absorbance at } 1470 \text{ cm}^{-1}(\text{PE Band})}, \quad (2)$$

$$\text{Ester/Aldhyde Index} = \frac{\text{Absorbance at } 1733 \text{ cm}^{-1}}{\text{Absorbance at } 1470 \text{ cm}^{-1}(\text{PE Band})}, \quad (3)$$

$$\text{Carboxylic Acid Index} = \frac{\text{Absorbance at } 1698 \text{ cm}^{-1}}{\text{Absorbance at } 1470 \text{ cm}^{-1}(\text{PE Band})}. \quad (4)$$

A sample of the compiled spectra of virgin/waste PE formulations, which were treated under accelerated weathered conditions are presented in Figs 1–3. Table 1 also shows the prominent and most noted band in PO polymer degradation and IR detection [29]–[31]. The band observed in the range between 700 to 760  $\text{cm}^{-1}$  corresponds to the rocking vibration associated to the crystallinity of the polymer [32]. The band between 1460–1480  $\text{cm}^{-1}$  relates to  $\text{CH}_2$  bending vibration, while the band between 2850–2950  $\text{cm}^{-1}$  associated with the asymmetric and symmetric CH stretching [32]. The band between 1000 and 1300  $\text{cm}^{-1}$  corresponds to the formulation of hydroperoxide  $\text{ROOH}$  [33]. In general, peroxide radical  $\text{ROO}^*$  is formed after macro free radical  $\text{R}^*$ , reacting with the ambient oxygen. The peroxide radical  $\text{ROO}^*$  can then separate hydrogen from the polymer, forming polymer centred free radical  $\text{P}^*$  and a hydroperoxide  $\text{ROOH}$  [28], [33]. As a result, the hydroperoxide  $\text{ROOH}$  degrades rapidly and forms highly reactive radicals that accelerate the oxidation process forming ketone, esters, and acids [33]. The major bands in Figs 1–3 were the spectra between 1700 and 1800  $\text{cm}^{-1}$ , which relates to the accumulation of carbonyl group and its products.



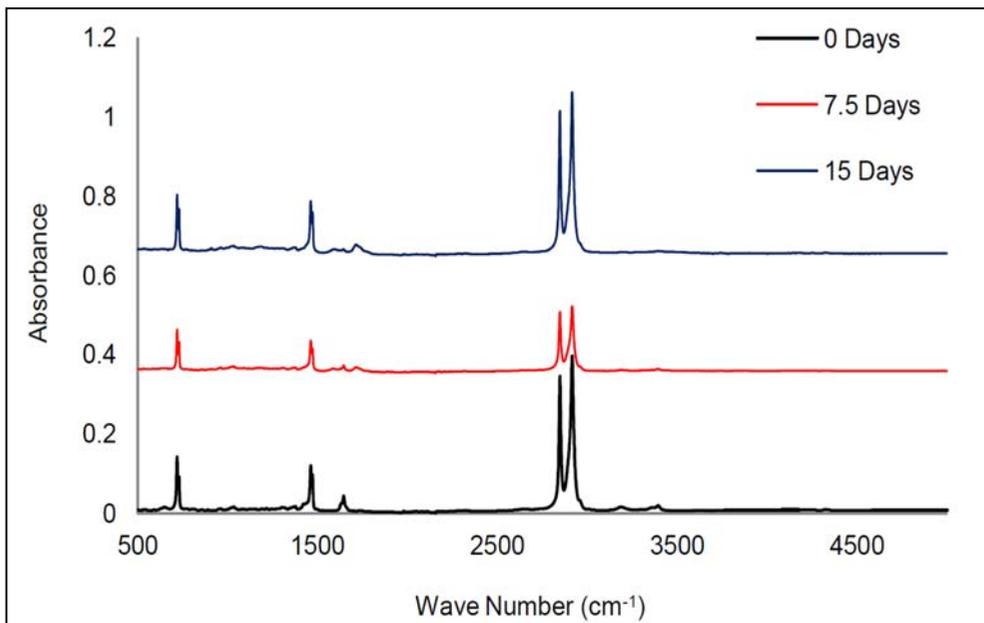


Figure 1: FTIR-ATR spectra for the 100 wt.% virgin linear low density PE samples exposed to accelerated weathering.

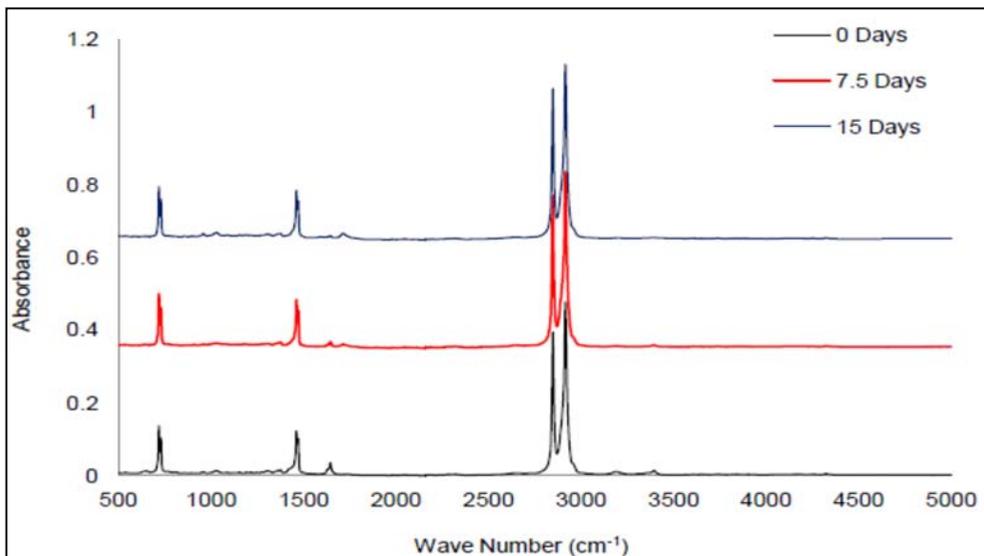


Figure 2: FTIR-ATR spectra for the 100 wt.% virgin linear low density PE and 25 wt.% PSW samples exposed to accelerated weathering.

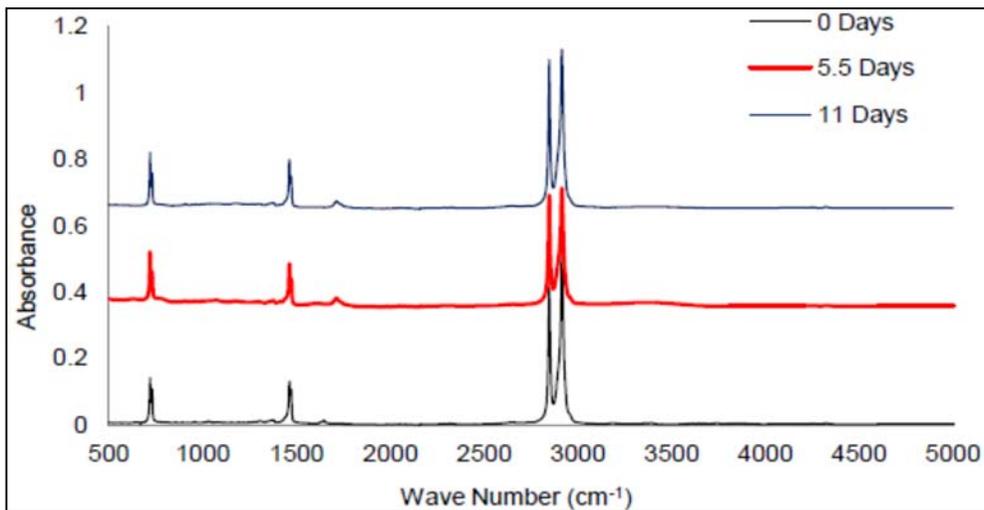


Figure 3: FTIR-ATR spectra for the 100 wt.% PSW samples exposed to accelerated weathering.

Table 1: Assignment of absorption bands of IR spectrum.

Wave number (cm <sup>-1</sup> )	Assigned Group/Product	Reference
1770	C=O Stretching (Lactone)	Gulmine et al. [29]
1733	C–O Stretching (Ester/Aldhyde)	
1714	C=O Stretching (Ketone)	
1470	PE Band	D'Amelia et al. [30]
1698	C=O Stretching (Carboxylic Acid)	Gulmine et al. [29]
730	CH <sub>2</sub> rocking vibration, crystalline	Matuana et al. [31]
718	CH <sub>2</sub> rocking vibration, amorphous	

The dominant absorption band at around 1714 cm<sup>-1</sup> corresponds to the C=O stretching vibration of terminal ketone [29], and it is more prominent with the increase in waste content by weight in the PE formulations (Figs 2, 3).

The ketone absorbance after 5.5 days of continuous exposure increased by approximately 46.7% between the 100 wt.% virgin polymer and the 100 wt.% PSW samples. In addition, with extended aging for each formulation, it was noted that the intensity of the absorption band increases. Additional bands for other carbonyl products start to appear such as ester/aldehyde band at 1733 cm<sup>-1</sup>, lactone band at 1770 cm<sup>-1</sup>, and carboxylic acid band at 1689 cm<sup>-1</sup> (see Fig. 4).

The CI is an indication for the oxidation degree for each sample, notably increasing with weathering duration and waste content of the samples. Therefore, the stability of the samples against oxidation is in the following order: 75V/25W > 50V/50W > 100V/0W > 25V/75W > 0V/100W. However, it is very important to mention that for 50% PE virgin content and above, the CI stabilized at the last stages of aging compared to samples with more waste content (> 50%). A good explanation for this oxidation rate stability could be due to the

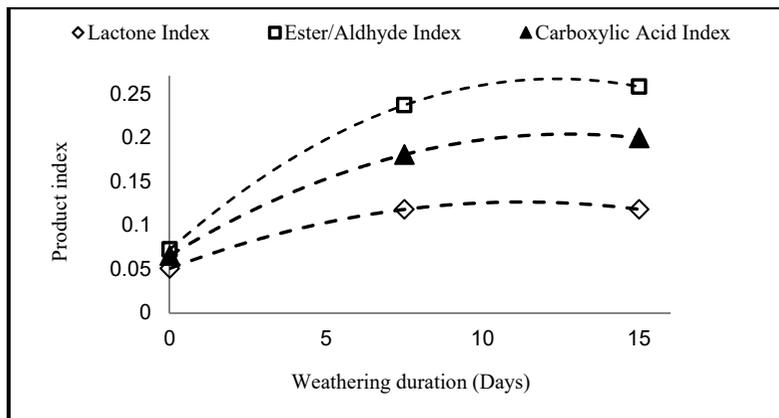


Figure 4: Variation of lactone index, ester/aldehyde index and carboxylic index for the 100 PE sample.

effects of the oxidation and cross-linking processes [29]. Samples with 50% virgin content and above favoured cross-linking over oxidation; therefore, carbonyl formation decreased and stabilized at last stages of aging [29]. For samples with 50% waste content and below, product indices homogeneously increase as weathering duration increases. Ester/aldehyde is the major oxidizing product (highest index) followed by carboxylic acid and lactone. Furthermore, the values of indices for all products increase as the virgin content in the samples increases from 50% to 100%.

Ester/aldehyde groups are being produced clearly, increasing from 0.12 (50V/50W) to 0.24 (100V/0W) for PE films with 50% waste content and below. The group also increases with aging time. This is an indication for the oxidation of virgin PE to accumulate carbonyl products at high values. Slight increase is noticed for Ester/aldehyde groups for PE films with 50% waste content and above (see figures above). The carboxylic acid and lactone indices increase with aging because of the oxidation of all samples [33]. Nevertheless, carboxylic acid index is always higher than lactone index for all PE samples. This is attributed to the fact that it is the final stable product of the oxidation process, which causes the accumulation of carboxylic acid [9]. Unlike the 50% waste content and below samples, heterogeneous trends are observed for 75% and 100% waste content in the samples. This can be clearly seen by the increase in products indices during the early stages of aging, while inconsistent pattern is noticed in the later stages. The values of indices for these samples also are lower than the values for 50% virgin content and above. This reflects the chemical degradation and oxidation that waste PE have already gone through during its lifetime; therefore, carbonyl products indices for these samples are lower and heterogeneous.

The degree of crystallinity is presented in Table 2 following the thermal characterisation conducted. The virgin LLDPE exhibited a similar degree of crystallinity as to the other materials tested in this work (34%). The structure of the materials tested in this work being semi-crystalline allows for the degradation within the polymers matrix due to weathering [34], [35]. Loss of chain ends due to oxygen permeability is denoted to be the main reason of the detected calorimetric behaviour [26]. The alteration in the mechanical integrity of the LLDPE samples studied in this work, namely in Young's modulus, strain and stress at rupture, indicate the loss of amorphous region at 7.5 days of exposure [16]. This points

Table 2: Degree of crystallinity (%) measurements and oxidation induction temperature (°C) from DSC thermal analysis.

Material (v/w wt.%)	Exposure duration (days)	Degree of crystallinity (%) 2nd heating cycle	Oxidation induction temp. (°C)
100/0	0	34	200.5
	7.5	40	196.6
	15	35	181
75/25	0	35	196.6
	7.5	35	195.2
	15	41	180
50/50	0	34	189.3
	5.5	32	185.6
	11	34	183.8
25/75	0	34	188.3
	5.5	34	187.8
	11	35	185
0/100	0	34	183.2
	5.5	33	175.7
	11	35	178.4

towards the excessive cross-linking and branching reactions that the polymer undergoes due to photo-degradation which are evident with the increase in crystallinity. The decrease in crystallinity at threshold was expected due to the increase in the evolution of chemicals from the termination reaction for the studied chemical mechanism associated with PE degradation [36].

The materials studied in this work were also subjected to thermogravimetric analysis and a sample of the results is shown in Fig. 5. The onset, mideste and maximum degradation temperatures showed very similar values in this investigation. This indicates that the waste content has no effect on the thermal characteristics of the material. It should also be noted that the content of 2 wt.% of polypropylene (PP) in the waste materials seems to be irrelevant on the thermogravimetric behaviour of the materials. This can be attributed to the fact that PP was in minimal weight content that did not warrant a dedicated step in the thermogram. The char content increased with the increasing waste content, which is related to the stabiliser content in the plastic products of west and number heating cycles the materials were subjected to. The thermal analysis has revealed valuable insight into the behaviour of PSW blends under different conditions that can be used as a guideline for future product development. The history of the PSW, being exposed to natural weathering at reclamation sight and processed with various heating loops, has influenced the specimens behaviour in this study. Thermo-mechanical degradation was an influential factor on the behaviour of the materials since the blends varied than the LLDPE samples in their characteristics [10], [15], [17].

#### 4 CONCLUSION

Blends constituting plastic solid waste (PSW) from polyolefin (PO) polymers origin were investigated in this work with the aim of determining the effect of the solid waste ratio on the materials properties. The focus was on the chemical properties and products evolved due to degradation which were investigated using infrared spectroscopy. Accelerated weathering was conducted in this work to determine the degradation profile of the samples. The waste



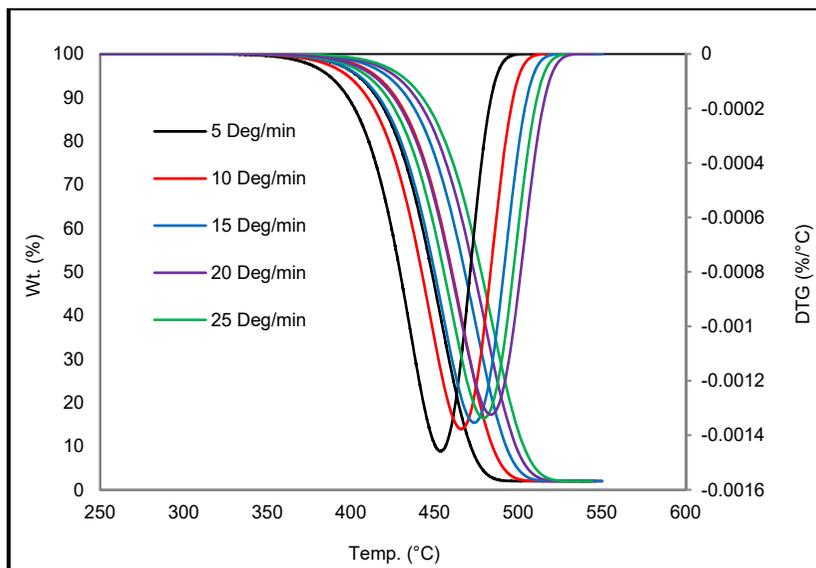


Figure 5: Thermogram of the virgin/waste (75/25) unexposed sample subjected to nitrogen weight loss analysis showing weight loss (%) (primary scale) and first derivative (secondary scale).

materials were incorporated using 25 to 100% by weight and the specimens were tested under two different weathering durations to showcase the degradation profile of the specimens. Thermal stability determine using thermogravimetry showed a somewhat comparable results in the measurements. This was attributed to the fact that the samples were all made of PO polymers and degrade in similar temperature ranges regardless of their past processing history. The crystallinity degree was noted to be similar before exposure to weathering tests. Minor changes with respect to the degree of crystallinity were noticed as weathering progresses which is a consequence of the change in crystal size. The 75/25 samples showed a change in crystallinity toward midpoint of weathering time resulting in a 41% degree of crystallinity from the 35% originally estimated. This was related to the immiscibility of the polymers and crystal size change due to oxidation. This is supported with the detected energy absorption curves noted in the DSC analysis. The structure of the LLDPE was also noted to be influential in the changes notable in the oxidation induction temperature (OIT) analysis carried out in this study. From the Thermogravimetry and DSC studies conducted, the 75/25 was noted to be most stable and can be used as a good material for product development applied for indoor and outdoor purposes. The morphological analysis has led to the detection of mosaic cracks in all directions that progressed as weathering took place. However, the reduction in the carbonyl index determined from the FTIR spectra of the 75/25 supports the claim of the relative stability of the blend. The work conducted in this study can benefit from density measurements and EDS analysis to support some of the experimental findings. This can lead to the development of a product from waste plastic material that reduced the burden of accumulation in developing and developed societies alike. The thermal degradation kinetics of the blend material should be assessed in the future to determine the materials recycling potential and viability via thermal and chemical means.

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